

हमारा विश्वास... हर एक विद्यार्थी है स्यास

NEET

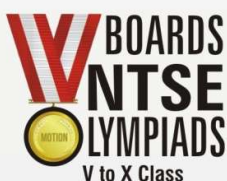
12th Sep. 2021

QUESTION PAPER
WITH
SOLUTION



CHEMISTRY

CODE – P2



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SECTION - A

51. The incorrect statement among the following is :

- (1) Actinoids are highly reactive metals, especially when finely divided.
- (2) Actinoid contraction is greater for element to element than Lanthanoid contraction.
- (3) Most of the trivalent Lanthanoid ions are colorless in the solid state.
- (4) Lanthanoids are good conductors of heat and electricity.

Sol. 3

Fact

52. Given below are two statements:

Statement I:

Aspirin and Paracetamol belong to the class of narcotic analgesics.

Statement II:

Morphine and Heroin are non-narcotic analgesics. In the light of the above statements, choose the correct answer from the options given below.

- (1) Statement I is incorrect but Statement II is true.
- (2) Both Statement I and Statement II are true.
- (3) Both Statement I and Statement II are false.
- (4) Statement I is correct but Statement II is false.

Sol. 3

Fact

53. Statement-I :

Acid strength increases in the order given as $\text{HF} \ll \text{HCl} \ll \text{HBr} \ll \text{HI}$.

Statement II :

As the size of the elements F, Cl, Br, I increases down the group, the bond strength of HF, HCl, HBr and HI decreases and so the acid strength increases.

In the light of the above statements, choose the correct answer from the options given below.

- (1) Statement I is incorrect but Statement II is true.
- (2) Both Statement I and Statement II are true.
- (3) Both Statement I and Statement II are false.
- (4) Statement I is correct but Statement II is false.

Sol. 2

In hydrides acidic strength increases on moving down the group as acid easily release H^+ and conjugate anion stabilise due to decrease in charge density

54. Which one among the following is the correct option for right relationship between C_p and C_v for one mole of ideal gas?

- (1) $C_v = RC_p$ (2) $C_p + C_v = R$ (3) $C_p - C_v = R$ (4) $C_p = RC_v$

Sol. 3

$C_p - C_v = R$



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55. The correct option for the number of body centred unit cells in all 14 types of Bravais lattice unit cells is:

- (1) 3 (2) 7 (3) 5 (4) 2

Sol. 1

Cubic	→ 1BCC] Total BCC 3
Tetragonal	→ 1BCC	
Orthorhombic	→ 1BCC	

56. Among the following alkaline earth metal halides, one which is covalent and soluble in organic solvents is :

- (1) Beryllium chloride (2) Calcium chloride
(3) Strontium chloride (4) Magnesium chloride

Sol. 1

Covalent character directly proportional to the polarisation &

Polarisation \propto charge on cation & anion

$$\text{Polarisation} \propto \frac{1}{\text{size of cation}}$$

Polarisation \propto size of anion

Here +ve = const

-ve = const

Size of anion = constant

Size of cation : $\text{Be}^{+2} < \text{Mg}^{+2} < \text{Ca}^{+2} < \text{Sr}^{+2}$

\therefore Polarisation $\text{Be}^{+2} > \text{Mg}^{+2} > \text{Ca}^{+2} > \text{Sr}^{+2}$

\therefore Covalent character $\text{BeCl}_2 > \text{MgCl}_2 > \text{CaCl}_2 > \text{SrCl}_2$

57. Tritium, a radioactive isotope of hydrogen, emits which of the following particles ?

- (1) Neutron (n) (2) Beta (β^-)
(3) Alpha (α) (4) Gamma (γ)

Sol. 2

58. The maximum temperature that can be achieved in blast furnace is :

- (1) upto 5000 K (2) upto 1200 K
(3) upto 2200 K (4) upto 1900 K

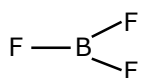
Sol. 3

Fact

59. BF_3 is planar and electron deficient compound. Hybridization and number of electrons around the central atom, respectively are :

- (1) sp^2 and 8 (2) sp^3 and 4
(3) sp^3 and 6 (4) sp^2 and 6

Sol. 4



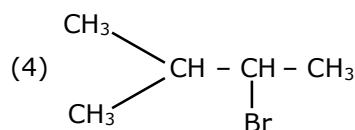
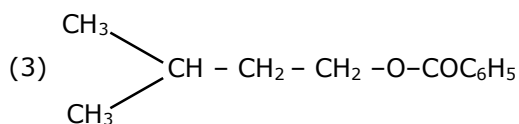
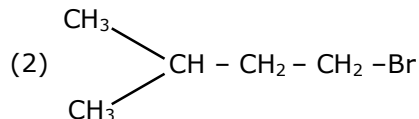
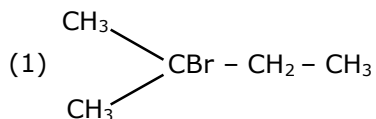
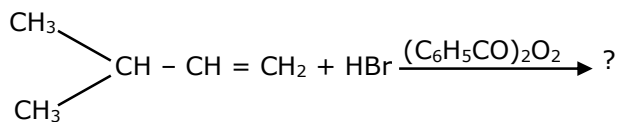
ℓp on CA = 0

σ -bond with boron = 3

\therefore Steric number = 3

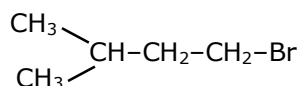
Hybridization = sp^2

60. The major product of the following chemical reaction is :



Sol. 2

Peroxide effect, major product according to anti markovnikov's rule (based on stability of free radical intermediate)



61. The molar conductance of NaCl, HCl and CH_3COONa at infinite dilution are 126.45, 426.16 and 91.0 $\text{S cm}^2 \text{mol}^{-1}$ respectively. The molar conductance of CH_3COOH at infinite dilution is. Choose the right option for your answer.

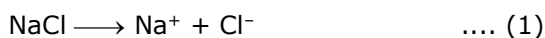
(1) 540.48 $\text{S cm}^2 \text{mol}^{-1}$

(2) 201.28 $\text{S cm}^2 \text{mol}^{-1}$

(3) 390.71 $\text{S cm}^2 \text{mol}^{-1}$

(4) 698.28 $\text{S cm}^2 \text{mol}^{-1}$

Sol. 3



$$(2) + (3) - (1)$$

$$426.16 + 91 - 126.45$$

$$= 390.71$$

62. A particular station of All India Radio, New Delhi, broadcasts on a frequency of 1.368 kHz (kilohertz). The wavelength of the electromagnetic radiation emitted by the transmitter is : [speed of light, $c = 3.0 \times 10^8 \text{ ms}^{-1}$]

(1) 21.92 cm

(2) 219.3 m

(3) 219.2 m

(4) 2192 m

Sol. 2

$$\gamma = \frac{c}{\lambda}$$

$$\lambda = \frac{c}{\gamma}$$

$$\lambda = \frac{3 \times 10^8}{1368 \times 10^3}$$

$$\lambda = 219.29 = 219.3 \text{ m}$$

63. Which of the following reactions is the metal displacement reaction ? Choose the right option.

- (1) $2\text{Pb}(\text{NO}_3)_2 \rightarrow 2\text{PbO} + 4\text{NO}_2 + \text{O}_2 \uparrow$ (2) $2\text{KClO}_3 \xrightarrow{\Delta} 2\text{KCl} + 3\text{O}_2$
 (3) $\text{Cr}_2\text{O}_3 + 2\text{Al} \xrightarrow{\Delta} \text{Al}_2\text{O}_3 + 2\text{Cr}$ (4) $\text{Fe} + 2\text{HCl} \rightarrow \text{FeCl}_2 + \text{H}_2 \uparrow$

Sol. 3



64. The right option for the statement. "Tyndall effect is exhibited by". is:

- (1) Urea solution (2) NaCl solution
 (3) Glucose solution (4) Starch solution

Sol. 4

Tyndall effect shown by starch solution

65. The compound which shows metamerism is:

- (1) $\text{C}_4\text{H}_{10}\text{O}$ (2) C_5H_{12} (3) $\text{C}_3\text{H}_8\text{O}$ (4) $\text{C}_3\text{H}_6\text{O}$

Sol. 1



This molecular formula is applicable for homologous series ether ($-\text{O}-$) a bivalent functional group and as we know ether with minimum four-C shows metamerism.

$\text{CH}_3-\text{CH}_2-\text{O}-\text{CH}_2-\text{CH}_3$ & $\text{CH}_3-\text{O}-\text{CH}_2-\text{CH}_2-\text{CH}_3$ are metamers.

66. Match List I with List-II.

List-I

- (a) PCl_5
 (b) SF_6
 (c) BrF_5
 (d) BF_3

List-II

- (i) Square pyramidal
 (ii) Trigonal planar
 (iii) Octahedral
 (iv) Trigonal bipyramidal

Choose the correct answer from the options given below :

- (1) (a)-(iv), (b)-(iii), (c)-(ii), (d)-(i)
 (2) (a)-(iv), (b)-(iii), (c)-(i), (d)-(ii)
 (3) (a)-(ii), (b)-(iii), (c)-(iv), (d)-(i)
 (4) (a)-(iii), (b)-(i), (c)-(iv), (d)-(ii)

Sol. 2

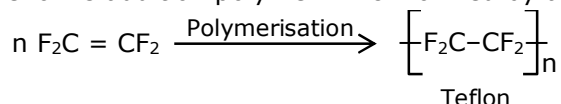
	SN(Stearic number)	
PCl_5	5	$\ell p = 0 \quad \sigma = 5$ Trigonal bipyramidal
SF_6	6	$\ell p = 0 \quad \sigma = 6$ octahedral
BrF_5	6	$\ell p = 1 \quad \sigma = 5$ square pyramidal
BF_3	3	$\ell p = 0 \quad \sigma = 3$ trigonal bipyramidal

67. Which one of the following polymers is prepared by addition polymerisation?

- (1) Dacron (2) Teflon (3) Nylon-66 (4) Novolac

Sol. 2

Teflon is addition polymer which formed by addition polymerisation of Tetra fluoroethene



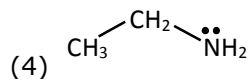
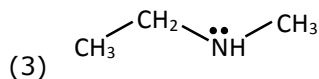
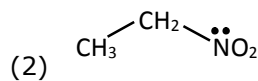
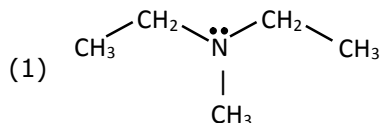
68. The RBC deficiency is deficiency disease of:

- (1) Vitamin B₂ (2) Vitamin B₁₂ (3) Vitamin B₆ (4) Vitamin B₁

Sol. 2

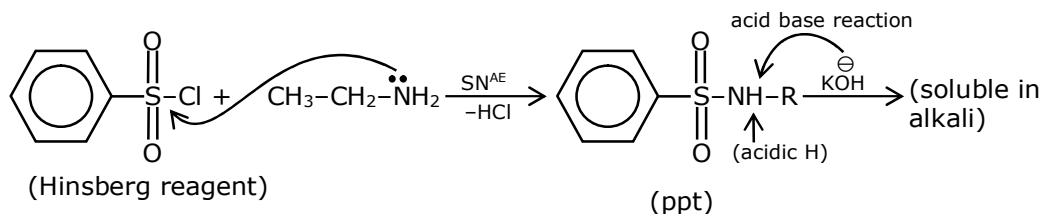
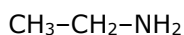
Vitamin B₁₂

69. Identify the compound that will react with Hinsberg's reagent to give a solid which dissolves in alkali.



Sol. 4

1° amine reacts with hinsberg reagent to give ppt. that is soluble in alkali.



70. Zr(Z=40) and Hf(Z=72) have similar atomic and ionic radii because of :

- (1) having similar chemical properties
(2) belonging to same group
(3) diagonal relationship
(4) lanthanoid contraction

Sol. 4

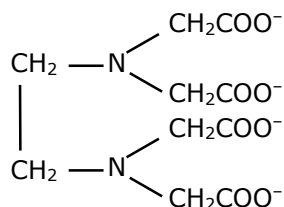
From IV A to II B

Due to lanthanoid contraction size of 4d & 5d series elements of same group are approx same.

71. Ethylene diaminetetraacetate (EDTA) ion is :

- (1) Tridentate ligand with three "N" donor atoms
(2) Hexadentate ligand with four "O" and two "N" donor atoms
(3) Unidentate ligand
(4) Bidentate ligand with two "N" donor atoms

Sol. 2

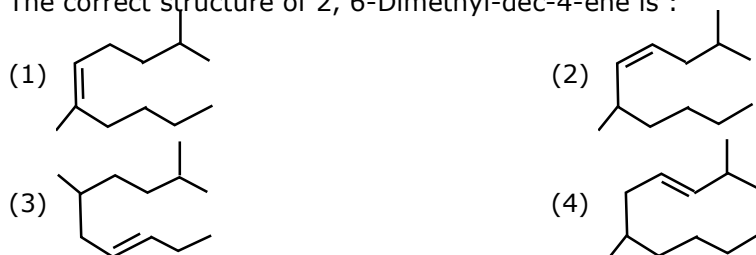


72. The major product formed in dehydrohalogenation reactions of 2-Bromo pentane is Pent-2-ene. This product formation is based on ?

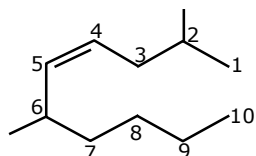
- (1) Huckel's Rule (2) Saytzeff's Rule
(3) Hund's Rule (4) Hofmann Rule

Sol. 1
HCP Structure
THV = 12
OHV = 6

78. The correct structure of 2, 6-Dimethyl-dec-4-ene is :

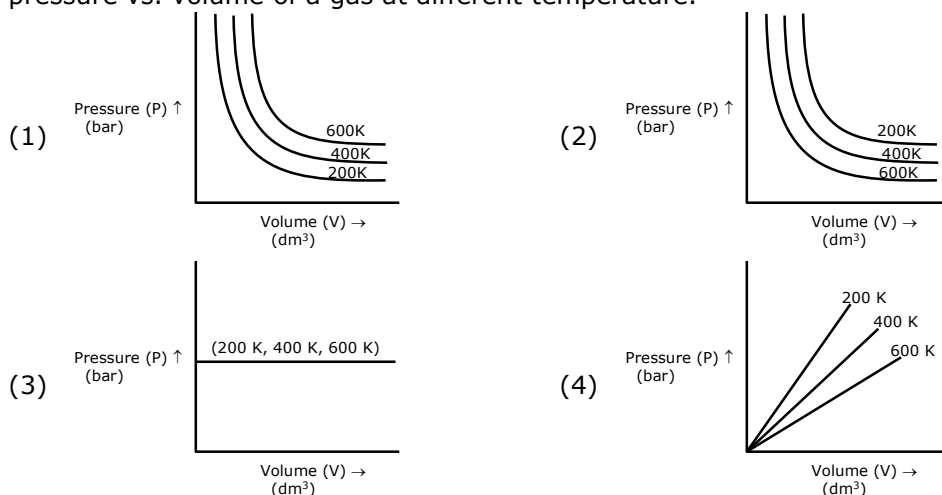


Sol. 2

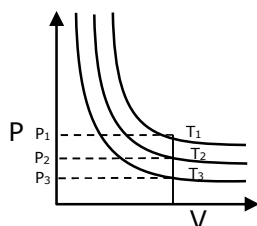


2, 6-Dimethyl dec-4-ene

79. Choose the correct option for graphical representation of Boyle's law, which shows a graph of pressure vs. volume of a gas at different temperature.



Sol. 1



$$PV = nRT$$

$$P \propto T$$

$$P_1 > P_2 > P_3$$

$$T_1 > T_2 > T_3$$

- 80.** The following solutions were prepared by dissolving 10g of glucose ($C_6H_{12}O_6$) in 250 ml of water (P_1), 10 g of urea (CH_4N_2O) in 250 ml of water (P_2) and 10 g of sucrose ($C_{12}H_{22}O_{11}$) in 250 ml of water (P_3). The right option for the decreasing order of osmotic pressure of these solution is:
 (1) $P_3 > P_1 > P_2$ (2) $P_2 > P_1 > P_3$ (3) $P_1 > P_2 > P_3$ (4) $P_2 > P_3 > P_1$

Sol.

2

$$\Pi = CRT$$

$$\Pi \propto C \Rightarrow \Pi \propto \frac{1}{M_w}$$

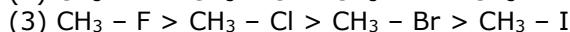
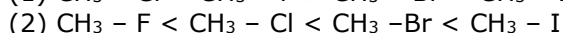
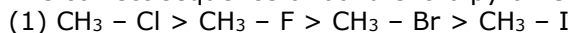
$$\text{Glucose} \Rightarrow M_{w1} = 180$$

$$\text{Urea} \Rightarrow M_{w2} = 60$$

$$\text{Sucrose} \Rightarrow M_{w3} = 342$$

$$P_2 > P_1 > P_3$$

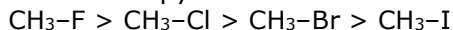
- 81.** The correct sequence of bond enthalpy of 'C-X' and is



Sol.

3

Bond enthalpy of C-X bond



Down the group with the increase in size of halogen in C-X bond, energy decreases.

- 82.** Dihedral angle of least stable conformer of ethane is :

(1) 0°

(2) 120°

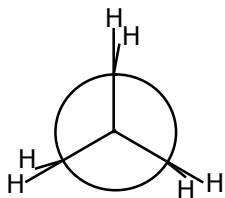
(3) 180°

(4) 60°

Sol.

1

Least stable conformer of ethane is eclipsed form



It's when dihedral angle is 0° .

- 83.** Noble gases are named because of their inertness towards reactivity. Identify an incorrect statement about them.

(1) Noble gases have large positive values of electron gain enthalpy.

(2) Noble gases are sparingly soluble in water.

(3) Noble gases have very high melting and boiling points.

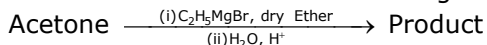
(4) Noble gases have weak dispersion forces.

Sol.

3

Fact

- 84.** What is the IUPAC name of the organic compound formed in the following chemical reaction?



(1) 2-methyl butan-2-ol

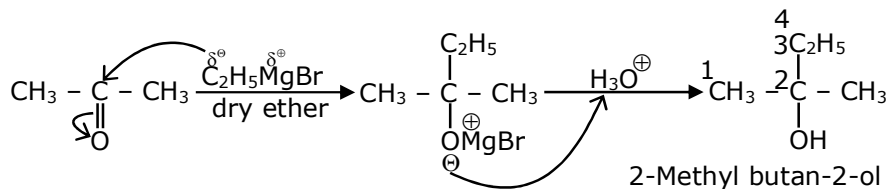
(2) 2-methyl propan-2-ol

(3) pentan-2-ol

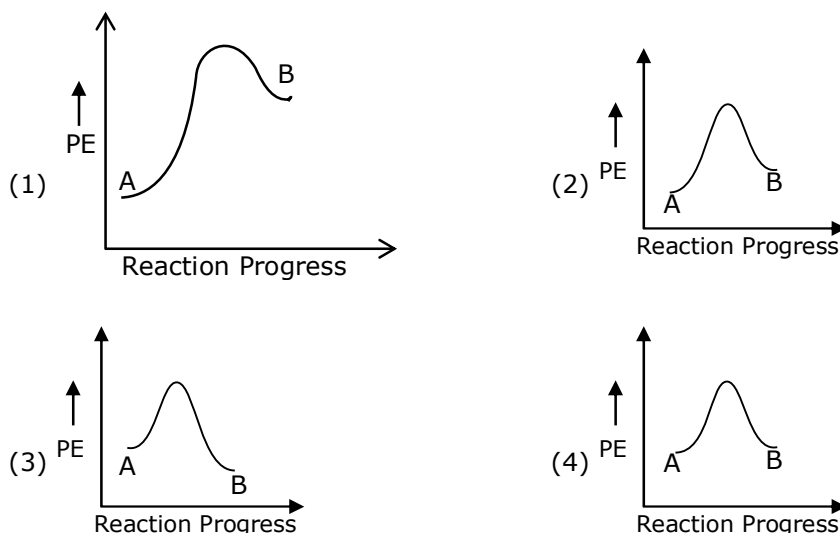
(4) pentan-3-ol

Sol.

1



85. For a reaction $A \rightarrow B$, enthalpy of reaction is -4.2 kJ mol^{-1} and enthalpy of activation is 9.6 kJ mol^{-1} . The correction potential energy profile for the reaction is shown in option.

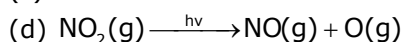
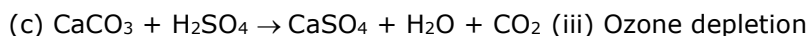
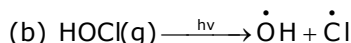
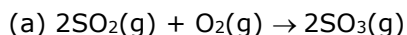


Sol. 3
 $\Delta H = (E_a)_f - (E_a)_b$
 $-4.2 = 9.6 - (E_a)_b$
 $(E_a)_b = 13.8$

SECTION - B

86. Match List-I with List-II.

List-I



List-II

(i) Acid rain

(ii) Smog

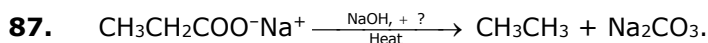
(iii) Ozone depletion

(iv) Tropospheric pollution

Choose the correct answer from the options given below.

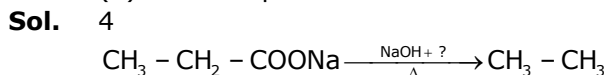
- (1) (a)-(iii), (b)-(ii), (c)-(iv), (d)-(i) (2) (a)-(i), (b)-(ii), (c)-(iii), (d)-(iv)
 (3) (a)-(ii), (b)-(iii), (c)-(iv), (d)-(i) (4) (a)-(iv), (b)-(iii), (c)-(i), (d)-(ii)

Sol. 4
 Fact



Consider the above reaction and identify the missing reagent/chemical.

- (1) DIBAL-H (2) B_2H_6
 (3) Red Phosphorus (4) CaO



It's decarboxylation in presence of sodalime so missing reagent is CaO.

(NaOH + CaO) sodalime

88. The correct option for the value of vapour pressure of a solution at 45°C with benzene to octane in molar ratio 3 : 2 is :

[At 45°C vapour pressure of benzene is 280 mm Hg and that of octane is 420 mm Hg. Assume Ideal gas]

- (1) 350 mm of Hg (2) 160 mm of Hg
 (3) 168 mm of Hg (4) 336 mm of Hg

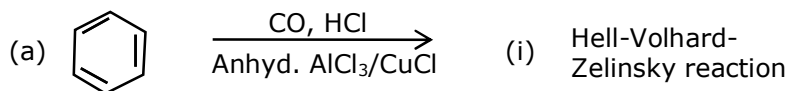
Sol. 4

$$\begin{aligned}P_s &= P_A^O x_A + P_B^O x_B \\&= 280 \times \frac{3}{5} + 420 \times \frac{2}{5} \\&= 56 \times 3 + 84 \times 2 \\&= 168 + 168 \\&= 336\end{aligned}$$

89. Match List-I with List-II

List-I

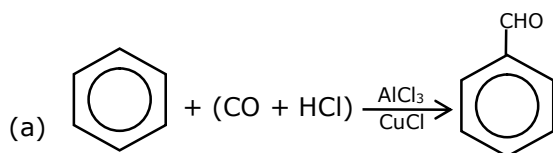
List-II



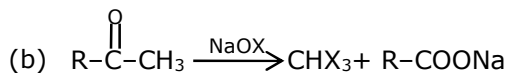
Choose the correct answer from the options given below :

- (1) (a)-(ii), (b)-(iii), (c)-(iv), (d)-(i)
- (2) (a)-(iv), (b)-(i), (c)-(ii), (d)-(iii)
- (3) (a)-(iii), (b)-(ii), (c)-(i), (d)-(iv)
- (4) (a)-(i), (b)-(iv), (c)-(iii), (d)-(ii)

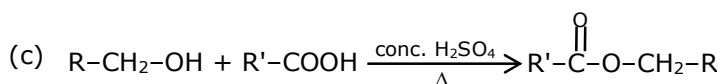
Sol. 1



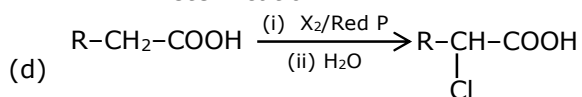
Gattermann Koch aldehyde synthesis



Haloform reaction

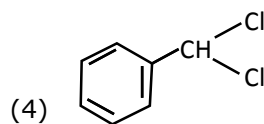
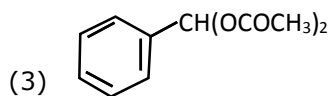
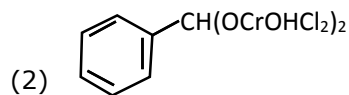
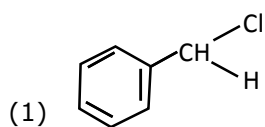
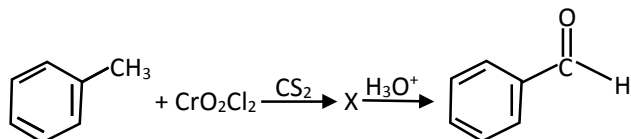


Esterification

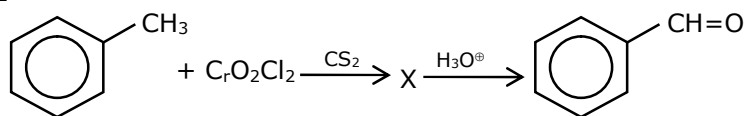


HVZ Reaction (Hell-volhard zelinsky Reaction)

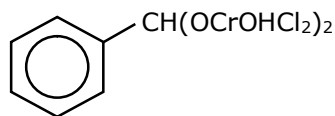
90. The intermediate compound 'X' in the following chemical reaction is :



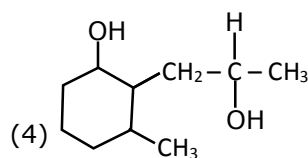
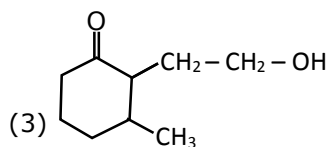
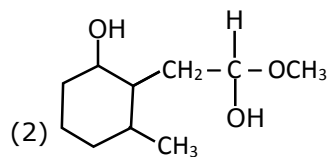
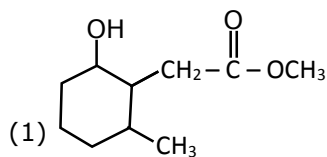
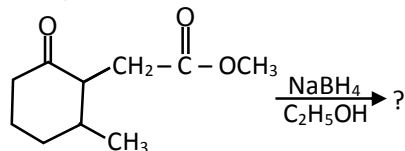
Sol. 2



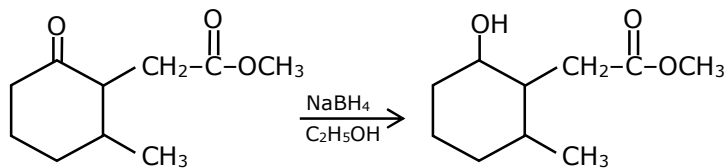
It's Etard Reaction "X" formed during the reaction



91. The product formed in the following chemical reaction is:



Sol. 1



NaBH₄ does not reduce ester group

92. The slope of Arrhenius Plot $\left(\ln K / \text{vs } \frac{1}{T}\right)$ of first order reaction is $-5 \times 10^3 \text{ K}$. The value of E_a of the reaction is. Choose the correct option for your answer.

[Given $R = 8.314 \text{ JK}^{-1}\text{mol}^{-1}$]

- (1) -83 kJ mol^{-1} (2) 41.5 kJ mol^{-1} (3) 83.0 kJ mol^{-1} (4) 166 kJ mol^{-1}

Sol. 2

$$\ln k = \ln A - \frac{E_a}{RT}$$

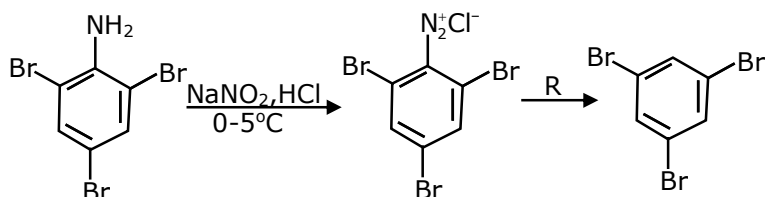
$$m = -\frac{E_a}{R}$$

$$\frac{-E_a}{R} = -5 \times 10^3$$

$$E_a = 5 \times 10^3 \times 8.314$$

$$E_a = 41.5$$

93. The reagent 'R' in the given sequence of chemical reaction is :



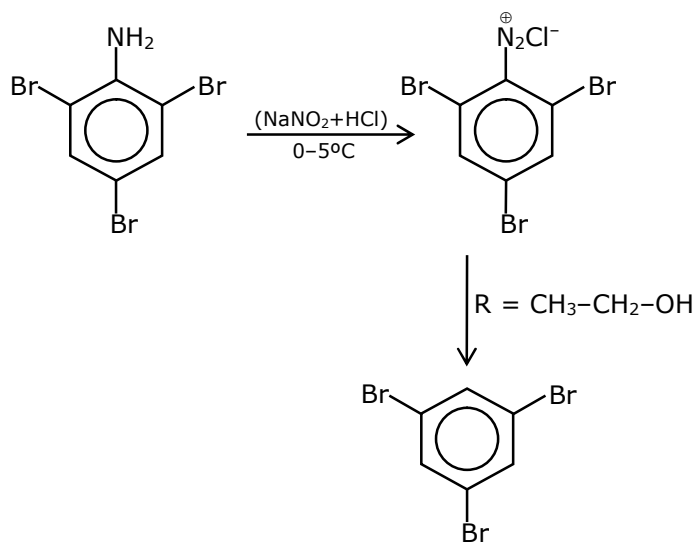
- (1) CuCN/KCN

- (2) H_2O

- (3) $\text{CH}_3\text{CH}_2\text{OH}$

- (4) HI

Sol. 3



94. For irreversible expansion of an ideal gas under isothermal condition, the correct option is :

- (1) $\Delta U \neq 0, \Delta S_{\text{total}} = 0$

- (2) $\Delta U = 0, \Delta S_{\text{total}} = 0$

- (3) $\Delta U \neq 0, \Delta S_{\text{total}} \neq 0$

- (4) $\Delta U = 0, \Delta S_{\text{total}} \neq 0$

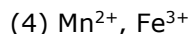
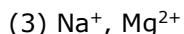
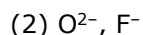
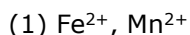
Sol. 4

$$\Delta S_{\text{total}} > 0$$

$$\Delta S_{\text{total}} \neq 0$$

$$\Delta U = 0$$

95. From the following pairs of ions which one is not an iso-electronic pair ?



Sol. 1

Not isoelectronic



96. The molar conductivity of 0.007 M acetic acid is $20 \text{ Scm}^2\text{mol}^{-1}$. What is the dissociation constant of acetic acid? Choose the correct option.

$$\left[\begin{array}{l} \Lambda_{\text{H}^+}^\circ = 350 \text{ Scm}^2 \text{mol}^{-1} \\ \Lambda_{\text{CH}_3\text{COO}^-}^\circ = 50 \text{ Scm}^2 \text{mol}^{-1} \end{array} \right]$$

(1) $2.50 \times 10^{-5} \text{ mol L}^{-1}$

(2) $1.75 \times 10^{-4} \text{ mol L}^{-1}$

(3) $2.50 \times 10^{-4} \text{ mol L}^{-1}$

(4) $1.75 \times 10^{-5} \text{ mol L}^{-1}$

Sol. 4

$$(\Lambda_m^\infty)_{\text{CH}_3\text{COOH}} = 50 + 350$$

$$= 400$$

$$\alpha = \frac{\Lambda_m}{\Lambda_m^\infty} = \frac{20}{400} = 0.05$$

$$= K_a = C\alpha^2$$

$$= 0.007 \times (0.05)^2$$

$$= 7 \times 10^{-3} \times 25 \times 10^{-4}$$

$$= 175 \times 10^{-7}$$

$$= 1.75 \times 10^{-5}$$

97. Which of the following molecules is non-polar in nature?

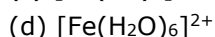
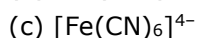
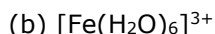
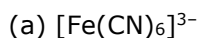


Sol. 4

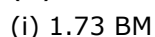
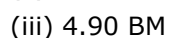
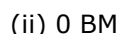
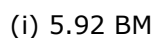
Fact

98. Match List – I with List – II.

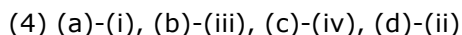
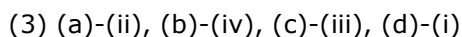
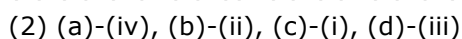
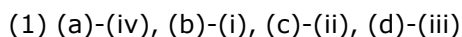
List – I



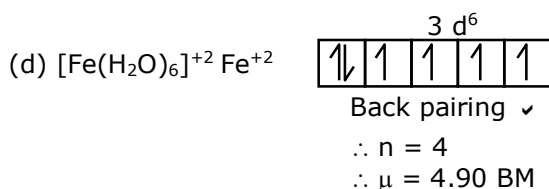
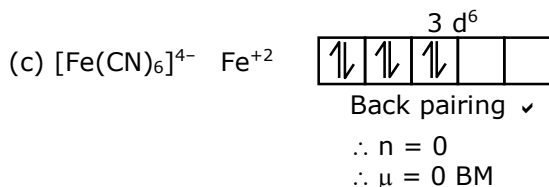
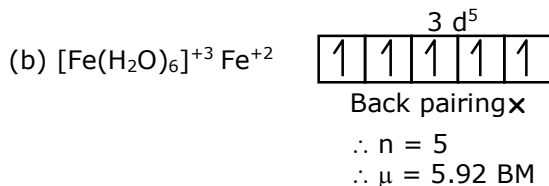
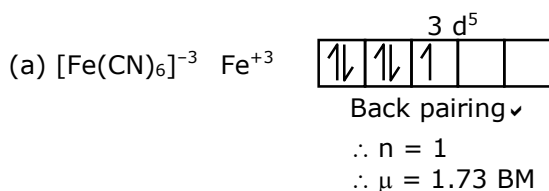
List – II



Choose the correct answer from the options given below.



Sol. 1



99. Choose the correct option for the total pressure (in atm.) in a mixture of 4g O_2 and 2g H_2 confined in a total volume of one litre at 0°C is :

[Given $R = 0.82 \text{ L atm mol}^{-1} \text{ K}^{-1}$, $T = 273 \text{ K}$]

- (1) 26.02 (2) 2.518
 (3) 2.602 (4) 25.18

Sol. 4

$$PV = nRT$$

$$P \times 1 = \left(\frac{4}{32} + \frac{2}{2} \right) \times 0.082 \times 273$$

$$P = 25.18 \text{ atm.}$$

100. In which one of the following arrangements the given sequence is not strictly according to the properties indicated against it?

- (1) $\text{CO}_2 < \text{SiO}_2 < \text{SnO}_2 < \text{PbO}_2$: Increasing oxidizing power
 (2) $\text{HF} < \text{HCl} < \text{HBr} < \text{HI}$: Increasing acidic strength
 (3) $\text{H}_2\text{O} < \text{H}_2\text{S} < \text{H}_2\text{Se} < \text{H}_2\text{Te}$: Increasing pK_a values
 (4) $\text{NH}_3 < \text{PH}_3 < \text{AsH}_3 < \text{SbH}_3$: Increasing acidic character

Sol. 3

Oxidising power : $\text{CO}_2 < \text{SiO}_2 < \text{SnO}_2 < \text{PbO}_2$

Acidic strength : $\text{HF} < \text{HCl} < \text{HBr} < \text{HI}$

pK_a : $\text{H}_2\text{O} > \text{H}_2\text{S} > \text{H}_2\text{Se} > \text{H}_2\text{Te}$

Acidic strength : $\propto \frac{1}{pK_a}$

Acidic strength : $\text{NH}_3 < \text{PH}_3 < \text{AsH}_3 < \text{SbH}_3$

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